

## **LEO the lion, says GER**

Lose electrons  
oxidization, gain  
electrons reduction

## **Acid**

Compound that gives  
off  $H^+$  ions in solution

## **Anion**

Ions with a negative charge.

## **Atom**

The smallest object that  
retains properties of an  
element. Composed of  
electrons and a nucleus  
(containing protons and  
neutrons).

## **Avogadro's Number**

Number representing the number of molecules in one (1) mole:  $6.022 \times 10^{23}$ .

## **Atomic Number**

Number of protons in an element.

## **Base**

Substance which gives off hydroxide ions ( $\text{OH}^-$ ) in solution.

## **catalyst**

Substance that speeds up a chemical process without actually changing the products of reaction.

## **cations**

Ion with a positive charge.

## **Charge**

Describes an object's ability to repel or attract other objects. Protons have positive charges while electrons have negative charges. Like charges repel each other while opposite charges, such as protons and electrons, attract one another.

## **chemical equation**

An expression of a fundamental change in the chemical substances.

## **compound**

Two or more atoms joined together chemically, with covalent or ionic bonds.

## **Covalent bond**

When two atoms share at least one pair of electrons.

## **density**

Mass per unit volume of a substance.

## **electron**

One of the parts of the atom having a negative charge. Indivisible particle with a charge of -1.

## **element**

Substance consisting of only one type of atom.

## **equilibrium**

When the reactants and products are in a constant ratio. The forward reaction and the reverse reactions occur at the same rate when a system is in equilibrium.

## **exothermic**

Process that gives off heat to the environment.

## **force**

An entity that when applied to a mass causes it to accelerate. Sir Isaac Newton's Second Law of Motion states: the magnitude of a force = mass \* acceleration.

## **half life**

The amount of time it takes for half an initial amount to disintegrate.

## **ionics bonds**

When two oppositely charged atoms share at least one pair of electrons but the electrons spend more time near one of the atoms than the other.

## **ion**

Removing or adding electrons to an atom creates an ion (a charged object very similar to an atom).

## **mass number**

The number of protons and neutrons in an atom.

## **mole**

A collection of  $6.022 \times 10^{23}$  number of objects. Usually used to mean molecules.